

Ionic Charges Chart

Cations

1+		2+		3+	
ammonium	NH_4^+	barium	Ba^{2+}	aluminum	Al^{3+}
cesium	Cs^+	beryllium	Be^{2+}	chromium(III)	Cr^{3+}
gold(I)	Au^+	cadmium	Cd^{2+}	cobalt(III)	Co^{3+}
hydrogen	H^+	calcium	Ca^{2+}	gold(III)	Au^{3+}
lead(I)	Pb^+	cobalt(II)	Co^{2+}	iron(III)	Fe^{3+}
lithium	Li^+	copper(II)	Cu^{2+}	manganese(III)	Mn^{3+}
potassium	K^+	iron(II)	Fe^{2+}		
silver	Ag^+	lead(II)	Pb^{2+}		
sodium	Na^+	magnesium	Mg^{2+}		
copper(I)	Cu^+	manganese(II)	Mn^{2+}		
		mercury(I)	Hg_2^{2+}	tin(IV)	Sn^{4+}
		mercury(II)	Hg^{2+}	nickel(IV)	Ni^{4+}
		nickel(II)	Ni^{2+}	lead(IV)	Pb^{4+}
		strontium	Sr^{2+}		
		zinc	Zn^{2+}		
		tin(II)	Sn^{2+}		

Roman numeral notation indicates charge of ion when element commonly forms more than one ion. For example, iron(II) has a 2+ charge; iron(III) a 3+ charge.

Anions

1-		2-		3-	
acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	cyanide	CN^-	carbonate	CO_3^{2-}
amide	NH_2^-	cyanate	OCN^-	chromate	CrO_4^{2-}
hydrogen carbonate (bicarbonate)	HCO_3^-	fluoride	F^-	dichromate	$\text{Cr}_2\text{O}_7^{2-}$
hydrogen sulfate (bisulfate)	HSO_4^-	hydride	H^-	oxide	O^{2-}
bisulfide	HS^-	hydroxide	OH^-	oxalate	$\text{C}_2\text{O}_4^{2-}$
bisulfite	HSO_3^-	hypochlorite	ClO^-	silicate	SiO_3^{2-}
bromate	BrO_3^-	iodate	IO_3^-	sulfate	SO_4^{2-}
bromide	Br^-	iodide	I^-	sulfide	S^{2-}
chlorate	ClO_3^-	nitrate	NO_3^-	sulfite	SO_3^{2-}
chlorite	ClO_2^-	nitrite	NO_2^-	tartrate	$\text{C}_4\text{H}_4\text{O}_6^{2-}$
chloride	Cl^-	perchlorate	ClO_4^-	tetraborate	$\text{B}_4\text{O}_7^{2-}$
		permanganate	MnO_4^-	thiosulfate	$\text{S}_2\text{O}_3^{2-}$
		thiocyanate	SCN^-		
				arsenate	AsO_4^{3-}
				arsenite	AsO_3^{3-}
				citrate	$\text{C}_6\text{H}_5\text{O}_7^{3-}$
				ferricyanide	$\text{Fe}(\text{CN})_6^{3-}$
				nitride	N^{3-}
				phosphate	PO_4^{3-}
				phosphite	PO_3^{3-}
				phosphide	P^{3-}

There are no common anions with a 4- charge.